***Year 13 Chemistry Class Activities – Week 10 and 11***

At the end of the class, students should be able to;

* **calculate** enthalpy change in simple systems involving either chemical change or phases changes using the following (Hess’s Law):

] Δr*H*° = ΣΔf*H*° (products) - ΣΔf*H*° (reactants)

Δr*H*° = Σ enthalpies of bonds broken – Σ enthalpies of bonds made

* **calculate** enthalpy change in complex systems involving either chemical change or phases changes using the following (Hess’s Law)

Δr*H*° = Σ Δf*H*° (products) - Σ Δf*H*° (reactants)

Δr*H*° = Σ enthalpies of bonds broken – Σ enthalpies of bonds made

***Student Activities.***

* Read the pdf attached more than two times



* Watch the link given below on youtube

Link 1: youtube.com/watch?v=OTm6JmeYX5g (Bond energy)

Link 2: youtube.com/watch?v=JRIm\_a2LDPM (Endothermic and Exothermic reactions)

* Do activity 4C and 4D on page 38 and 42
* Cross check your answers with the solution below.

Solutions to activity 4C and 4D

(check next page)



For any queries, contact me on messenger; mosese vereti